Chemistry F105X: General Chemistry I, 4.0 Credits
Fall Semester, 2019

Instructor:	Dr. William A. Howard	
Office:	Reichardt Building Room 190	
Laboratory	Reichardt Building Room 191	
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Email:	wahoward@alaska.edu	
Course Type:	In person	
Lecture Meetings:	Mon., Wed., Fri. 1:00 to 2:00 PM	Reichardt Room 201
Office Hours:	Mon., Wed., Fri. 2:00 to 3:00 PM	

Prerequisites: The student should be placed in Math F151X or higher and in Writing F111X or higher. Furthermore, the student should have had at least one year of chemistry in high school or should have received a B- or higher in Chemistry F103X.

Required Course Materials: The following materials are used this course:

1. Two semester access to Smartwork, our online homework system. Smartwork will be used in both General Chemistry F105X and in General Chemistry F106X, and so students will need two semester access. Included with Smartwork is an electronic text book.
2. "Preparing for your ACS Examination in General Chemistry" by Eubanks and Eubanks. ISBN 978-0970804204.
3. Non-programmable calculator. (A programmable calculator is one that has graphing functions and can show graphs on the display. Programmable calculators are not permitted in this class.) Students will need their calculators when taking exams and doing homework problems.
4. The text book is …
   Chemistry: An Atoms-Focused Approach, 2nd edition, by Gilbert, Kirss, Foster, & Bretz
   Published by Norton
   Copyright: 2018
   ISBN-13: 9780393284218

There is no text book for the lab. Rather, handouts for each experiment will be posted on Blackboard.

Catalog Description:

CHEM F105X General Chemistry I (n)
4 Credits
This course is an introduction to general chemistry and explores topics to a much greater depth than preparatory courses. Topics include: measurement, energy and matter, periodic trends, chemical composition, chemical reactions, solutions, bond theory, gases, thermodynamics, problem-solving (applied mathematics), and special topics. Special Note: CHEM F105X-F106X, together with their laboratory components, constitute the

Course Calendar:

General Chemistry F105X: Schedule of Lectures			
Fall, 2019			**Professor William A. Howard**
Week No.	**Date**	**Reading**	**Classroom Lecture / Activity**
1	8/26/19	No reading	Introduction to Course. Discuss syllabus.
	8/28/19	Sections 1.1 – 1.3	Atomic Theory, Classifications of Matter
	8/30/19	Sections 1.4 – 1.7	States of Matter, Precision & Accuracy, Significant Figure
2	9/2/19		**NO CLASS – LABOR DAY**
	9/4/19	Sections 1.8 – 1.9	Unit Conversions, Dimensional Analysis, Statistics
	9/6/19	Sections 2.1 – 2.2	Atomic Structure, Nuclei
3	9/9/19	Sections 2.3 – 2.4	Periodic Table, Atomic Masses
	9/11/19	Sections 2.5 – 2.6	Moles, Molar Masses, Mass Spectroscopy
	9/13/19	Sections 3.1 – 3.3	Light, Atomic Spectra, Photons, Photoelectric Effect
4	9/16/19	Section 3.4	Bohr Model of Hydrogen
	9/18/19	Sections 3.5 – 3.6	Electrons as Waves, Quantum Numbers
	9/20/19	Sections 3.6 – 3.8	Quantum Numbers, Atomic Orbitals, Auf Bau Principle
5	9/23/19	Sect. 3.9 – 3.12	Configurations of Ions, Radii, Ionization Energies, More!
	9/25/19		**PRACTICE TEST 1 (COVERS CHAPTERS 1 – 3)**
	9/27/19	Sections 4.1 – 4.2	Chemical Bonds, Lattice Energy, Formulas, More!
6	9/30/19	Sections 4.2 – 4.3	Nomenclature, Lewis Structures
	10/2/19	Sections 4.4 – 4.6	Resonance, Covalent Bonds, Electronegativity
	10/4/19		**REAL EXAM 1 (COVERS CHAPTERS 1 – 3)**
7	10/7/19	Sections 4.7 – 4.9	Formal Charges, Exceptions to Octet Rule
	10/9/19	Sections 5.1 – 5.2	VSEPR Theory
	10/11/19	Sections 5.3 – 5.4	Molecular Polarity, Hybridization of Atomic Orbitals
8	10/14/19	Sections 5.5 – 5.7	Simple Organics, Chirality, Molecular Orbital Theory
	10/16/19	Section 5.7	Molecular Orbital Theory
	10/18/19	Sections 6.1 – 6.2	Intermolecular Forces
9	10/21/19	Sections 6.3 – 6.5	Solubility, Phase Diagrams, Water
	10/23/19		**PRACTICE TEST 2 (COVERS CHAPTERS 4 – 6)**
	10/25/19	Sections 7.1 – 7.3	Balancing Chemical Reactions, Stoichiometry
10	10/28/19	Sections 7.4 – 7.5	Empirical Formulas, Molecular Formulas
	10/30/19	Sections 7.6 – 7.7	Combustion Analysis, Limiting Reagents, % Yield
	11/1/19		**REAL EXAM 2 (COVERS CHAPTERS 4 – 6)**
11	11/4/19	Sections 8.1 – 8.3	Solutions, Dilutions, Electrolytes
	11/6/19	Sections 8.4 – 8.5	Acid-Base Chemistry, Net Ionic Equations
	11/8/19	Section 8.6	Redox Reactions, Oxidation Numbers
12	11/11/19	Sections 8.7 – 8.8	Titrations, Ion Exchange
	11/13/19	Sections 9.1 – 9.4	Energy, Heat, Work, Enthalpy

	11/15/19	Sections 9.5 – 9.6	Calorimetry, Hess' Law
13	11/18/19	Section 9.7	Calculating Changes in Enthalpy
	11/20/19	Sections 9.8 – 9.9	Born-Haber Cycles and More Enthalpy Calculations
	11/22/19	Sect. 10.1 – 10.2	Gases, Kinetic Theory, Graham's Law
14	11/25/19	Sect. 10.3 – 10.7	Pressure, Gas Laws
	11/27/19		**NO CLASS – THANKSGIVING BREAK**
	11/29/19		**NO CLASS – THANKSGIVING BREAK**
15	12/2/19	Sect. 10.8 – 10.10	Gas Stoichiometry, Dalton's Law of Partial Pressures
	12/4/19		**PRACTICE TEST 3 (COVERS CHAPTERS 7 – 10)**
	12/6/19		**REAL EXAM 3 (COVERS CHAPTERS 7 – 10)**
16	**12/13/19**		**Final Examination, 1:00 – 3:00 PM**

standard one year engineering and science major general chemistry course. Students must be enrolled in both CHEM F105X and CHEM F105L to receive full credit.
Prerequisites: Placement in MATH F151X; placement in WRTG F111X; or a B- or better in CHEM F103X; or permission of instructor and department.
Corequisite: CHEM F105L.
Attributes: UAF GER Natural Science Req
Lecture + Lab + Other: 3 + 3 + 0

Course Overview: General Chemistry F105X, a 4.0 credit course, is the first semester of a two semester series in general chemistry which describes a wide variety of microscopic and macroscopic chemical phenomena. We will cover chapters 1 – 10 of the Gilbert text, according to the schedule accompanying this syllabus. Chem F105X is a "depth" core course. Your attendance at lecture, Mondays, Wednesdays, Fridays 1:00 – 2:00 PM in Reichardt room 201, is expected.
General Chemistry F106X is the second-semester sequel to the Chem F105X course, and chapters 11 – 19 and 21 of the text book will be covered. At the end of Chem F106X, a standardized final examination will be given. This final examination will test for some concepts covered in Chem F105X.

Course Goals and Student Learning Outcomes: As a result of the General Chemistry F105X experience, students will become familiar with and practice the scientific method and learn basic skills in laboratory practices, in general chemistry, and in performing chemical calculations. Students will also learn how to test ideas experimentally and will learn the relationship between science and public policy. Some key events in the history of chemistry will also be learned, and students will become familiar with naming compounds and the "language" of chemistry.
Learning the topics in this course will be accomplished by reading the text book, attending lecture, solving problems, taking exams, and actively participating in the laboratory component of the course. Learning the scientific method will be accomplished in the laboratory by carrying out lab experiments and writing about the results in lab reports.

Instructor's Expectations of Students: There is a VAST amount of material covered in the text book. This class moves fairly quickly through this material, and it is simply impossible for Prof. Howard to cover all of the material fully in the lectures! Students are expected to read the text book faithfully according to the schedule shown in the table on pages 2 and 3 of this syllabus.
Furthermore, students must attend class faithfully. Important class announcements are usually given at the beginning of each lecture, and students can save themselves a lot of trouble simply be being aware of what's going on in class.
Students should bring paper and a non-programmable calculator to class every day, in order to solve in-class problems that are a part of the lectures.
Students are expected to attend and participate in lab each week. Any student who completes fewer than 8 lab reports will fail this course, regardless of his or her exam grades and other scores.

Grades: There are some conditions that, if met, will result in an "automatic grade." These conditions are:
1) The final exam in this course will be a standardized, national final exam, prepared by the American Chemical Society. If a student obtains a final exam score that is in the 90^{th} percentile nationally (or higher), then that student will receive a final letter grade of A for this course, provided that the student has completed at least 8 lab reports.
2) If a student obtains a final exam score that is equal to or higher than the 80^{th} percentile nationally but lower than the 90^{th} percentile, then that student will receive a final letter grade of B for this course, provided that the student has completed at least 8 lab reports. (If this student would have obtained an A in this course without this special condition, then the student will be given an A as the final letter grade.)
3) If a student does not take the final exam and has no legitimate reason (sickness, death in the family, etc.) for missing the exam, then that student will receive a letter grade of F for this course – regardless of all other homework and exam scores.
4) If a student completes fewer than 8 lab reports, then that student's final letter grade for this course will be an F, regardless of all other factors.

If a student does not meet any of these special conditions, then that student's grade will be determined by either (1) the student's final percentage score (described below), or (2) a class grading curve (also described below).

Percentage Score: This class will have a total of …

Attendance	36 points
Smart Work Homework	408 points
In-Class Practice Problems	30 points
3 Practice Exams	120 points
3 Real Exams	300 points
Take Back 3 Real Exams	6 points
Lab Grade	100 points
Final Exam	100 points
Total Points	1100 points

At the end of the semester, a student's total score will be calculated by adding all the scores of the individual assignments. The student's total score will be divided by 1100, and the result will be multiplied by 100% to get the student's percentage score. The grades will be determined as follows:

Percentage score $\geq 90\%$	Final letter grade = A
$80\% \leq$ percentage score $< 90\%$	Final letter grade = B
$70\% \leq$ percentage score $< 80\%$	Final letter grade = C
$60\% \leq$ percentage score $< 70\%$	Final letter grade = D
Percentage score $< 60\%$	Final letter grade = F

Grading Curve: Let a be the average number of points for the students in this class, and let σ be the standard deviation. If x = the total number of points that a student has, then the final letter grade is calculated according to the following curve:

$x \geq a + (1.5)\sigma$	A
$a + 1\sigma \leq x \leq a + (1.5)\sigma$	B
$a + \frac{1}{2}\sigma \leq x \leq a + 1\sigma$	C+
$a - \frac{1}{2}\sigma \leq x \leq a + \frac{1}{2}\sigma$	C
$a - 1\sigma \leq x \leq a - \frac{1}{2}\sigma$	C–
$a - (1.5)\sigma \leq x \leq a - 1\sigma$	D
$x \leq a - (1.5)\sigma$	F

The class average (a) and the standard deviation (σ) will change every day, and so Prof. Howard will update the class before each lecture on the current values so that students can see how they are doing.

At the end of the semester, Prof. Howard will determine a letter grade for each student by BOTH methods (the percentage score and the grading curve). The final letter grade will be the higher of the two grades.

Opportunities to Earn Points: Opportunities to earn points include, but are not necessarily limited to, the options shown below.

1. *Attendance.*
   An attendance sheet will be passed around the class every day. Each student should print his or her initials on this attendance sheet. Prof. Howard will give each student 1 point per class-day if that student's name is recorded on the attendance sheet.
   If a student misses class for a good reason, then that student should show Dr. Howard a written excuse (from a doctor or a coach, etc.). For such an excused absence, the student will be given 1 point for attendance.
   If a student misses class and does not have a written excuse for missing class, then Dr. Howard will still give the student 1 point for attendance, up to 3 points maximum. In other words, students are allowed to have 3 unexcused absences

from class with no effect on grades; after the 3rd unexcused absence however, no points will be given for attendance if a student misses class.

2. *Smartwork 5 problems.*
All homework assignments are online. This class will use the Smartwork 5 online homework system, which can be accessed by Blackboard. Go to the chem 105 class on Blackboard and select the "course materials" link at the left side of the screen; there will be a folder for the Smartwork 5 homework assignments. The point value, due date and time, and estimated time for completion for each chapter homework assignment are summarized in the following table.

Homework Assignment	**Due Date & Time**	**Point Value**	**Estimated Time for Completion**
Getting Started	Sept. 4, 11:00 AM	57 points	1 hour, 30 minutes
Chapter 1	Sept. 6, 11:00 AM	31 points	1 hour, 14 minutes
Chapter 2	Sept. 13, 11:00 AM	37 points	1 hour, 23 minutes
Chapter 3	Sept. 25, 11:00 AM	32 points	1 hour, 59 minutes
Chapter 4	Oct. 9, 11:00 AM	50 points	2 hours, 38 minutes
Chapter 5	Oct. 18, 11:00 AM	31 points	1 hour, 7 minutes
Chapter 6	Oct. 23, 11:00 AM	27 points	1 hour, 26 minutes
Chapter 7	Nov. 4, 11:00 AM	51 points	3 hours, 2 minutes
Chapter 8	Nov. 13, 11:00 AM	45 points	2 hours, 43 minutes
Chapter 9	Nov. 22, 11:00 AM	26 points	2 hours, 46 minutes
Chapter 10	Dec. 4, 11:00 AM	21 points	1 hour, 27 minutes

- Late homework is accepted, but a penalty of 10% loss in points is applied for each day the homework is late; after 10 days, the homework is worth 0 points. For the chapter 10 homework only, a penalty of 33% loss in points is applied per day late.
- Students are given an infinite number of attempts on each homework problem.
- After an assignment is due, a student can re-open the assignment for practice problems in order to prepare for the exams.
- No homework will be accepted for any reason after 5 PM on Saturday, December 7.
- Occasionally, students experience problems using Smart Work 5. For example, a student may type in a right answer, but Smart Work 5 will count his answer as wrong. Or, perhaps a student cannot open Smart Work 5 on his particular laptop for some unknown reason. If a student experiences any "electronic" problems using Smart Work 5, the student must fill out a Help Desk ticket at https://helpdesk.wwnorton.com/external/helpdeskrequest.aspx?tid=11118. Reporting an electronic problem to Prof. Howard will not result in a solution to the problem.

3. *Practice Exams*

Except for exam 3, approximately 7 to 10 days before each regular in-class exam, there will be a practice exam given in class. (For exam 3, the practice test will be given on the class day before the real exam is given.) Each practice exam is worth 40 points and shall consist of 20 multiple choice problems – very similar to the real exam.

After taking the practice exam, the exam will be graded and returned to the students. An answer key will be posted on Blackboard, so that students can see which problems they missed and what the correct answers are. Then, each student will be given an opportunity to do point recovery, which is described below:

Point Recovery: A student may do point recovery work in order to win back the points lost on the practice exam. Point recovery work may be done by either of 2 ways: (1) by email, or (2) in-person.

Email Point Recovery: A student may solve a problem that he missed on the practice exam on a sheet of paper, and then scan the page and email the scanned page to Dr. Howard. If a student does point recovery by email, then the following rules will apply:

(a) Only 1 problem is to be shown on one page of paper. If a problem is very short, then 2 problems may be shown on the same page of paper. No more than 2 problems per page will be accepted.

(b) No typed work will be accepted for any reason. All point recovery work must be hand-written. The print must be LARGE and legible.

(c) For math problems, all work must be shown. For conceptual problems, a full explanation for how to solve the problem must be given, AND the page number in the text book where the explanation can be found must be given for reference. No points will be given if work is missing or if the explanation is incomplete or incorrect.

(d) For all point recovery work that is completed correctly BEFORE 5:00 PM on the day before the real exam is given, 2 points will be given for each problem. Thus, if a student does point recovery for all the missed problems on the practice exam before the day prior to the real exam, then the student will obtain a perfect score of 40 points for the practice exam! For all point recovery completed correctly after the day prior to the real exam, then the student can get a maximum of 1 point per problem.

(e) No point recovery work is accepted for any reason after 5 PM on Saturday, December 7.

In-Person Point Recovery: If a student visits Prof. Howard in his office and shows Prof. Howard how to solve the problem that he or she missed on the practice exam, then Prof. Howard will give that student some points. (The number of points given depends on when the point recovery work is done, as described in point d above.) In person point recovery is done on a first-come, first-serve basis; in other words, if 2 students want to do in-person point recovery at the same time, the student who arrived at Prof. Howard's office first is given

first priority. Rules c, d, and e (described above) apply to the in-person point recovery as well as to the email point recovery.

4. *Regular, in-class exams.*
   There will be 3 regular, in-class exams. Each is closed-book and closed-notes and will consist of 20 multiple choice problems. Each exam is worth 100 points.
   ** If a student misses an exam for a legitimate reason, that student may make up the exam at a later time. The student should contact Prof. Howard in order to set up a time to take the exam.

5. *Receiving your Exam Answer Sheet*
   After grading exams, the answer sheets will be returned to the students. A student will receive 2 points simply for taking his or her answer sheet. An answer key to the real exam will be posted on Blackboard, and students will be able to see which problems they missed and what the correct answers are.

6. *Other possible assignments*
   A student may propose assignments not listed in this syllabus in order to get more points. The student and Prof. Howard must agree on the educational value of the assignment, the number of points to be given, and how the assignment will be graded. Then, the student will be given an opportunity to persuade the class to accept the assignment. If the majority of the class votes to accept the assignment, then all students will have equal opportunity to complete this assignment. Again, the student MUST meet with Prof. Howard before beginning the work; if the student simply does an extra assignment without approval from Prof. Howard and the rest of the class, then no points will be given for that work.

7. *Laboratory Experiments*
   The purpose of the lab is to do hands-on investigation. We expect you to gain skills in scientific reasoning, experimental design, and use of chemicals and laboratory apparatus. The labs are conducted by graduate and upper division undergraduate teaching assistants who will have specific office hours. Lab reports will be given on Blackboard each week. Eleven experiments are scheduled for the semester, but the laboratory portion of your grade (100 points) will be based upon the average of your best ten lab grades.
   All students enrolled in Chem F105X must attend laboratory. Students completing reports for fewer than eight labs will fail the course, even if they have passing exam grades. You must attend lab prior to writing a lab report! There are no make-up labs scheduled during the semester. However, a student may make up a lab later in that same week if there is available space in another lab and with the TA's permission. If you have special scheduling problems or if you miss more than one lab for an acceptable reason, please discuss alternative plans with Emily Reiter, our general chemistry lab coordinator. Ms. Reiter's email address is e.reiter@alaska.edu and her phone number is 474-6748.

Laboratory reports are due one week after a lab is completed. Late lab reports will be accepted provided that the student has a good excuse for not turning in the lab report on time. The last report of the semester will not be accepted late.
Students MUST arrive in the lab on time. If a student is 15 minutes late, then that student will NOT be permitted to take the lab that day.
Cell phones and contact lenses are NOT permitted in the lab.
Please see the portion of the UAF Honor Code reproduced in the "Ethical Considerations" section below. Do not believe any rumors that it is acceptable to make up data or to use the work of another student (other than a lab partner in a collaborative experiment) as the basis of a lab report.

Week	**Dates**	**Lecture Topic**	**Lab Experiment**
1	8/26 to 8/30	Atomic Theory	NO LAB
2	9/2 to 9/6	Math with Significant Figures	NO LAB, math review handout
3	9/9 to 9/13	Physical Properties of Atoms	Chemical Health & Safety
4	9/16 to 9/20	Atomic Structure	Glassware, lab practice, Excel
5	9/23 to 9/27	Periodic Trends	Intro to Spectroscopy
6	9/30 to 10/4	Valence Bond Theory	Identify unknown solids
7	10/7 to 10/11	VSEPR Theory	Lewis structures, molecular modeling
8	10/14 to 10/18	Molecular Orbital Theory	Intermolecular forces
9	10/21 to 10/25	Phase Diagrams	Determination of sugar content
10	10/28 to 11/1	Formulas & Stoichiometry	Fun with stoichiometry
11	11/4 to 11/8	Aqueous Chemistry	Aqueous chemical reactions
12	11/11 to 11/15	Calorimetry	Copper cycle redox chemistry
13	11/18 to 11/22	Enthalpy & Gases	Introduction to gas laws
14	11/25 to 11/29	THANKSGIVING	NO LAB
15	12/2 to 12/6	Gas Stoichiometry	Review Sessions

8. *Final exam*
   The final examination will be given on Friday, December 13, 1:00 – 3:00 PM in room 201. The final examination will cover everything taught in General Chemistry F105X. The final exam is a standardized multiple-choice exam, prepared by the American Chemical Society, and worth 100 points. There are 70 questions on the final exam.
   All students will be required to take the final examination in order to pass this course.

Saturday, December 7 at 5 PM: All work (homework assignments, lab reports, exams, anything and everything) MUST be turned in before 5 PM on Saturday, December 7. Absolutely nothing is accepted for any reason after 5 PM on the last day of classes.

Help: Extra help is available, in case the student is having difficulty with Chem F105X. The student should visit Dr. Howard during normal office hours or make an appointment to see Dr. Howard by email or phone.
All teaching assistants (TAs) will have regular office hours also, and a student may see any TA for help with either lab or lecture. The TAs will see students in room 170 of the Reichardt Building. A schedule of office hours for TAs will be posted at the web site for the Chemistry Learning Center (http://www.uaf.edu/chem/clc/).

More Help:

1. UAF Department of Communication's Speaking Center
   907-474-5470
   speak@uaf.edu
2. UAF English Department's Writing Center (8th Floor, Gruening Building)
   907-474-5314
3. CTC Learning Center (604 Barnette Street)
   907-455-2860
4. UAF Math Lab (Room 305, Chapman Building)
   Department of Mathematics & Statistics 907-474-7332
   uafmathstatlab@gmail.com

Disabilities and Special Accommodations: Students with documented disabilities who may need reasonable academic accommodations must provide documentation of the disability to Disability Services in room 208 Whitaker Building, 474-5655, TTY 474-1827. Disability Services will then notify Prof. Howard in writing of the disability and will advise on how the student should be tested and which accommodations should be made.

Veteran Support Services: Walter Crary (wecrary@alaska.edu) is the Veterans Service Officer at the Veterans Resource Center (111 Eielson Building, 474-2475). Fairbanks Vet Center 456-4238. VA Community Based Outpatient Clinic at Ft. Wainwright is 361-6370.

Blackboard: All course documents will be placed on Blackboard. Grades will be updated on Blackboard regularly, and Prof. Howard will regularly inform the students of the class average and standard deviation. Each student will be able to see only his or her own scores and will NOT be able to see the scores of other students.
Freshman progress grades will NOT be posted on Blackboard; rather, these letter grades will be posted on UAOnline. Likewise, final letter grades are posted on UAOnline only, not Blackboard.

Instructor Withdrawals: The instructor reserves the right to withdraw any student from class for any of the following reasons:

(1) The student has not participated significantly as of November 1;
(2) The student has missed more than three labs as of November 1;
(3) The student does not work safely in the laboratory, and his or her continued presence in this course may pose a safety hazard for the student or for others.

Freshman Progress Reports: Freshman progress reports will be based on all graded materials, excluding lab grades, up to and including October 5.

Incompletes: The letter "I" (Incomplete) is a temporary grade used to indicate that the student has satisfactorily completed (C or better) the majority of work in a course but for personal reasons beyond the student's control, such as sickness, has not been able to complete the course during the regular semester. Negligence or indifference are not acceptable reasons for an "I" grade.

Ethical Considerations: The Chemistry "Department Policy on Cheating" is this: "*Any student caught cheating will be assigned a course grade of F. The student's academic advisor will be notified of this failing grade and the student will not be allowed to drop the course*."
Some examples of cheating include, but is not limited to, the following: (1) using cell phones during exams or quizzes; (2) plagiarism; (3) using programmable calculators during exams; (4) copying another student's answer while taking an exam or a quiz; and (5) unauthorized collaborations on lab projects. During exams and quizzes, electronic devices such as pagers, mp3 players, ipods, earbuds, text messengers, gameboys, etc. must be turned off.
As a UAF student, you are subject to UAF's Honor Code:
"Students will not collaborate on any quizzes, in-class exams, or take-home exams that will contribute to their grade in a course, unless permission is granted by the instructor of the course. Only those materials permitted by the instructor may be used to assist in quizzes and examinations.
Students will not represent the work of others as their own. A student will attribute the source of information not original with himself or herself (direct quotes or paraphrases) in compositions, theses, and other reports.
No work submitted for one course may be submitted for credit in another course without the explicit approval of both instructors.
Violations of the Honor Code will result in a failing grade for the assignment and, ordinarily, for the course in which the violation occurred. Moreover, violation of the Honor Code may result in suspension or expulsion."

Important Dates:

- Labor Day — Sept. 2
- Last day for registration or for adding classes — Sept. 6
- Last day for withdrawals with class not appearing on record — Sept. 6
- Last day to apply for Fall 2019 graduation — Oct. 15
- Early Progress Reports due — Oct. 7
- Last day for withdrawals with student receiving "W" — Nov. 1
- Thanksgiving Break — Nov. 27 – 29
- Last day of class — Dec. 7

Student Protections and Services Statement: Every qualified student is welcome in my classroom. As needed, I am happy to work with you, disability services, veterans'

services, rural student services, etc to find reasonable accommodations. Students at this university are protected against sexual harassment and discrimination (Title IX), and minors have additional protections. As required, if I notice or am informed of certain types of misconduct, then I am required to report it to the appropriate authorities. For more information on your rights as a student and the resources available to you to resolve problems, please go the following site: www.uaf.edu/handbook/